

Unit 6 – Equations and Stoichiometry

C3.4A Use the terms endothermic and exothermic correctly to describe chemical reactions in the laboratory.

C3.4c Write chemical equations including the heat term as a part of equation or using ΔH notation.

C5.2A Balance simple chemical equations applying the conservation of matter.

C5.2B Distinguish between chemical and physical changes in terms of the properties of the reactants and products.

C5.2d Calculate the mass of a particular compound formed from the masses of starting materials.

C5.2d.a I can use stoichiometry/dimensional analysis to calculate the amount of reactants and products in a given balanced chemical equation

C5.2e Identify the limiting reagent when given the masses of more than one reactant.

C5.2f Predict volumes of product gases using initial volumes of gases at the same temperature and pressure.

C5.6b Predict single replacement reactions.